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phase line) is characteristic for materials which have a higher coordination number in the liquid than in the solid phase (H<sub>2</sub>O, Bi, Si, Ge). Metals, for example have an opposite slope of the solid/liquid phase line, and the liquid phase will condense under pressure to a solid phase.

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Introduction to Materials Science, Chapter 9, Phase Diagrams University of Tennessee, Dept. of Materials Science and Engineering 6 A system is at equilibrium if at constant temperature, pressure and composition the system is stable, not changing with time. Equilibrium is the state that is achieved given sufficient time.

### **Chapter 9 Total - University of Tennessee**

Bergeron, C.G. and Risbud, S.H. (1984) Introduction to Phase Equilibria in Ceramics, The American Ceramic Society, Westerville, OH. This should be available in every ceramics laboratory. This should be available in every ceramics laboratory.

### **Equilibrium Phase Diagrams | SpringerLink**

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"Phase Diagram for Mercury up to 67 GPa and 500 K." Phys Rev B 48 (1993); 14009-14012.

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